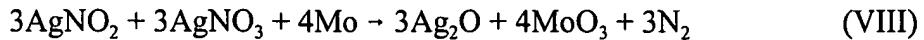


In the Specification:

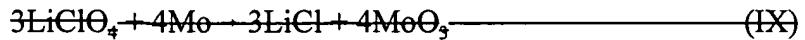
Amend as follows at page 18:

Example 3.



*C1*  
A comelt of equimolar amounts of  $\text{AgNO}_2$  and  $\text{AgNO}_3$  was mixed with a stoichiometric amount of Mo metal in accordance with equation VIII, i.e., 34.1% by weight  $\text{AgNO}_2$ , 37.6% by weight  $\text{AgNO}_3$ , and 28.3% by weight Mo. An autoignition temperature of  $131 \pm 2^\circ\text{C}$  was determined for the composition using DSC.

Example 4.



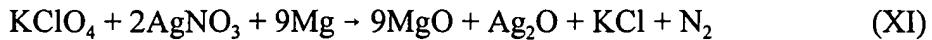
*C2*  
Lithium perchlorate,  $\text{LiClO}_4$ , was mixed with a stoichiometric amount of Mo in accordance with equation IX, i.e., 45.4% by weight  $\text{LiClO}_4$  and 54.6% by weight Mo. An autoignition temperature of  $147 \pm 2^\circ\text{C}$  was determined for the composition using DSC.

Example 5.



*C2*  
 $\text{AgNO}_3$  was mixed with a stoichiometric amount of magnesium, Mg, metal in accordance with equation X, i.e., 73.7% by weight  $\text{AgNO}_3$  and 26.3% by weight Mg. An autoignition temperature of  $157 \pm 2^\circ\text{C}$  was determined for the composition using DSC.

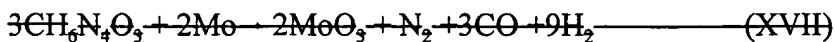
Example 6.



Amend as follows at page 21:

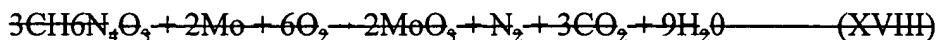
metal in accordance with equation XVI, i.e., 20.5% by weight  $\text{NaNO}_3$ , 41.0% by weight  $\text{AgNO}_3$  and 38.5% by weight Mo. The composition autoignited at  $217\pm 2^\circ\text{C}$  by DSC analysis.

Example 12:



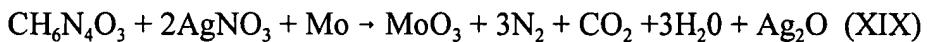
Guanidine nitrate,  $\text{CH}_6\text{N}_4\text{O}_3$ , was mixed with a stoichiometric amount of Mo in accordance with equation XVII, i.e., 60.4% by weight  $\text{CH}_6\text{N}_4\text{O}_3$  and 39.6% by weight Mo. The composition autoignited at  $230\pm 2^\circ\text{C}$  by DSC analysis.

This is an underoxidized reaction which leaves some products in an incompletely oxidized state. If there is an external source of oxygen the reaction proceeds according to equation XVIII.



This composition points out the utility of using organic nitrates in autoignition reactions.

Example 13.



A 1:2 ratio of guanidine nitrate to  $\text{AgNO}_3$  was mixed with a stoichiometric amount of Mo in accordance with equation XIX, i.e., 21.9% by weight  $\text{CH}_6\text{N}_4\text{O}_3$ , 60.9%  $\text{AgNO}_3$  and 17.2% by weight Mo. The composition autoignited at  $172\pm 2^\circ\text{C}$  (by DSC).

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Concl.

This composition is also an example of organic nitrates in autoignition reactions. However, this composition is fully oxidized, and, therefore, requires no external source of oxygen.

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